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## of Reaction Chapter 11 Chemical Reactions

Chapter 11: Properties of Reactions. An oxidation number is a positive or negative number that is assigned to an atom to indicate its degree of oxidation or reduction. The term oxidation state is often used interchangeably with oxidation

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number. A partial electron transfer is a shift in the electron density near an atom as a result of a change in the other atoms to which it is covalently bonded.

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## Chemical Reactions

answer key. coefficient. a whole number that appears before a formula in an equation. spectator ion. a particle not directly involved in a chemical reaction. combustion reaction. a reaction in which oxygen reacts with another substance, often producing light or heat. reactant.

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representation of a chemical reaction with  
reactants on the left, products on the right,  
and an arrow separating the two skeleton  
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\_\_\_\_\_ is a representation of a chemical reaction with reactants on the left, products on the right, and an arrow separating the two.

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precipitate. bubbles. signs of a chemical



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reaction. subscript. Solid compound produced from 2 aqueous solutions during a chem. Gas given off during a chemical reaction; one of the signs of. change in heat/light, formation of a gas (bubbles), solid prec.

chemistry chapter 11 chemical reactions

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Flashcards and ...

Write a balanced chemical equation for each reaction. Use the necessary symbols from Table 11.1 to describe the reaction completely.

a. Bubbling chlorine gas through a solution of potassium iodide gives elemental iodine and a solution of potassium chloride.

b. Bubbles of

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hydrogen gas and aqueous iron (III) chloride are produced when metallic

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Section 11.1 Assessment. Describe the steps in writing a balanced chemical equation. Write the skeleton equation for

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the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride.

$$\text{CuS (s)} + \text{O}_2\text{(g)} \rightarrow \text{Cu (s)} + \text{SO}_2\text{(g)}$$

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Students will relate how the Law of  
Conservation of Matter is used to  
determine chemical composition in  
compounds and chemical reactions. SC2.a.

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Identify and balance the following types of chemical equations: Synthesis, Decomposition, Single Replacement, Double Replacement, and Combustion  
SC2.b.

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cleavage reactions: A chemical reaction that has the general form  $AB \rightarrow A + B$ . Cleavage reactions are the reverse of condensation reactions.), and oxidation-reduction reactions A chemical reaction that exhibits a change in the oxidation states of one or more elements in the reactants that has the general form

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oxidant + reductant  $\rightarrow$  reduced oxidant + oxidized reductant.

Chapter 11.6: Types of Chemical  
Reactions - Chemistry ...

Chapter 11 Chemical Reactions 113

SECTION 11.1 DESCRIBING

CHEMICAL REACTIONS... A chemical

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reaction occurs when one or more \_\_\_\_\_  
change into one Chapter a I to ChemICal  
reaCtIons 4.1 Chemical Reactions and  
Chemical Equations 127 A chemical  
equation is a shorthand description of a  
chemical reaction.

Chapter 11 Chemical Reactions Practice

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Problems Answer Key ...

In the chemical reaction given, the reactants are sodium and water and the (formed) products are hydrogen and sodium hydroxide.

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Chemical Reactions ...

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## Chemical Reactions

As shown in Figure 11.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano. Heat, light, and gas are produced as a large pile of fluffy green chromium (III) oxide forms. We can describe this reaction with

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a chemical equation An expression that gives the identities and quantities of the substances in a chemical reaction.

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A chemical reaction means a rearrangement of the configuration to

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acquire the electronic configuration of the nearest stable or noble gas by transferring or sharing the electrons. Question 5. What is a valence electron?

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challenging the brain to think enlarged and faster can be undergone by some ways.

Experiencing, listening to the other experience, adventuring, studying, training, and more practical endeavors

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may encourage you to improve.

## Chemical Reactions Chapter 11 Test B Answer Key

Chapter 11 Chemical reactions What is a chemical reaction? its a reaction that changes the make up of an element and also creates a new substance. a chemical

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## Chemical Reactions

reaction occurs when compounds, elements, or atoms meet.

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