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oxygen in water.  $(16/18) * 100 = 88.9\%$   
Percentage by mass of element in a compound = (mass of element in 1 mol of  
compound ÷ molar mass of compound ) x 100 %  
(after completing lab #1) Find the % of  
carbon in sodium bicarbonate (NaHCO<sub>3</sub> ).

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water. Percentage by mass of element in a  
compound = (mass of element in 1 mol of  
compound ÷ molar mass of compound ) x 100 %.  
(after completing lab #1) Find the % of  
carbon in sodium bicarbonate (NaHCO<sub>3</sub> ).  
Example Problem: Find the % by mass of oxygen  
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Lab results: 1 doz grains of rice = \_\_\_\_\_g

(Use this fact as a conversion

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Note Taking Guide: Episode 703 Name\_\_\_\_\_

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Molecular Formula • represents the \_\_\_\_\_

number of \_\_\_\_\_ of each

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6ee0c2810b9aabdc290fa7f1cecf6975