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Dilution Problems, Chemistry, Molarity

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Chemistry 11 Calculations Practice Test #
1 . 13. Chemistry 11 Calculations Practice
Test # 2 . Molarity Worksheet # 1 . 1. 15.8

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g of KCl is dissolved in 225 mL of water.
Calculate the molarity.

~~Molarity Worksheet # 1~~

molarity of the diluted solution be? (0.75
M)(250 mL) = M₂ (295 mL) M₂ = (0.75
M)(250 mL) = 0.64 M (295 mL) 2) If
water is added to 175 mL of a 0.45 M

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KOH solution until the volume is 250 mL,
what will the molarity of the diluted
solution be? $(0.45 \text{ M})(175 \text{ mL}) = M_2$
 (250 mL)
 $M_2 = (0.45 \text{ M})(175 \text{ mL}) = 0.32$
 $M (250 \text{ mL})$

~~Dilutions Worksheet W 329 Everett
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Dilution Problems Worksheet 1. How do you prepare a 250.-ml of a 2.35 M HF dilution from a 15.0 M stock solution? 2. If 455-ml of 6.0 M HNO_3 is used to make a 2.5 L dilution, what is the molarity of the dilution? 3. If 65.5 ml of HCl stock solution is used to make 450.-ml of a 0.675 M HCl dilution, what is

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Molarity WS - HN KEY. Name: Part 1:
Molarity $M = \frac{\text{moles of solute}}{\text{Volume of Solution (L)}}$ Date: Molarity and Dilutions
Practice - = Mols Block: 1. 2. 3. 4. 5. 6.
What is the molarity of a 0.30 liter

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~~Answers~~ solution containing 0.50 moles of sodium chloride. Calculate the molarity of 0.289 moles of Iron (III) Chloride, FeCl_3 , dissolved in 120 of 1000 FL.

~~Molarity WS HN KEY~~

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~~Liberty Union High School District /~~

~~Overview~~

$M_1 V_1 = M_2 V_2$ (1.71 M) (25.0 mL) =
M₂ (65.0 mL) M₂ = 0.658 M. M = mol/L
= (25.0/40.0) / (0.325) = 1.92 mol/L. g =
(M) (L) (FW) = (0.400) ((0.225) (119) =
10.7 g. (25.0g) (1 mol/101 g)
(1000mL/0.650 mol) = 381 mL. Zn

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$(\text{NO}_3)_2 \text{AlCl}_3 \text{CuAc}_2$. 2 mol $\text{Ca}(\text{OH})_2 =$
mol HBr_2 (g/74) = (3.00) (0.0500) 5.55 g
 $\text{Ca}(\text{OH})_2$.

~~Molarity 1 (Worksheet) Chemistry
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Concentrations And Dilutions Answer
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Dilutions Worksheet - Solutions 1) If I add 25 mL of water to 125 mL of a 0.15 M NaOH solution, what will the molarity of the diluted solution be? $M_1V_1 = M_2V_2$
 $(0.15 \text{ M})(125 \text{ mL}) = x (150 \text{ mL})$ $x = 0.125$

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Answers
M 2) If I add water to 100 mL of a 0.15 M NaOH solution until the final volume is 150 mL, what will the molarity of the diluted solution be? $M_1V_1 = M_2V_2$

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Dilutions Worksheet □ Solutions 1) If I have 340 mL of a 0.5 M NaBr solution,

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Answers
what will the concentration be if I add 560 mL more water to it? 0.19 M (the final volume is 900 mL, set up the equation from that) 2) If I dilute 250 mL of 0.10 M lithium acetate solution to a volume of 750 mL, what will the concentration of this solution be?

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Answers, Dilutions work name key,
Solutions work 2 molarity and dilution
problems answers.

~~Dilution Worksheets - Kiddy Math~~

This worksheet features 5 molarity
problems ($M = \text{mol/L}$) with conversions
from grams to moles and milliliters to

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Answers and 7 dilutions problems using
 $M_1V_1=M_2V_2$. ANSWER KEY
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~~Answers~~
This worksheet and quiz will let you practice the following skills: Reading comprehension - ensure that you draw the most important information from the related how to calculate molarity and ...

~~Quiz & Worksheet - How to Calculate
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Answers of the stock solution? 4.

~~Molarity and Dilutions Worksheet~~
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□ molarity: the number of moles in a liter (volume), $M = \text{mol/L}$ □ equation for dilutions: $M_1V_1 = M_2V_2$, the concentration (or molarity) x volume of

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~~Answers~~ your original solution = the new concentration x new volume o In this case, the number of moles stays the same but the volume changes.

~~Molarity and Serial Dilutions Teacher Handout~~

This worksheet provides many examples

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~~Answers~~ for students to practice calculations involving Molarity & Molality. A complete answer key is provided at the end. This worksheet can be used in any Chemistry class, regardless of the students' ability level.

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