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The mole is the most common unit used to express the quantity of a chemical substance. For all solids, liquids, and gases, you can convert mass to moles (or moles to mass). For gases, you should memorize the following conversion of volume to moles (or moles to volume): at 0°C and 1 atmosphere pressure (known as standard temperature and pressure or STP ), one mole of any gas occupies approximately 22.4 liters.

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Chapter 10 Study Guide The Mole Answer Key

Mole Chapter 11: The Mole [on23j7zx8ml0] You know that the mole is defined as the number of representative particles, or carbon-12 atoms, in exactly 12 g of pure carbon-12. Thus, the mass of one mole of carbon-12 atoms is 12 g. What about other elements? Chapter 11: The Mole - Neshaminy Page 10/36

Chapter 11 The Mole

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Chapter 11 The Mole Study Guide Answers

14 1 mole = 6.02 x 1023 molecules (covalent) 1 mole = 6.02 x 1023 formula units (ionic) HOW MANY PARTICLES. 1 mole = 6.02 x 1023 atoms (monoatomic element) 1 mole = molar mass

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